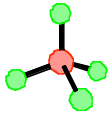

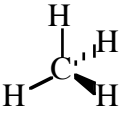
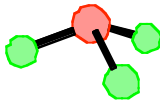

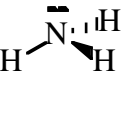
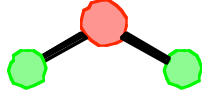

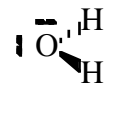
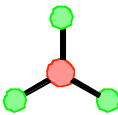
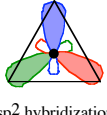
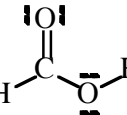
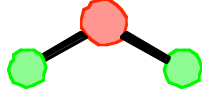
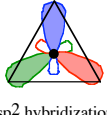
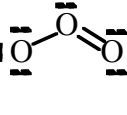


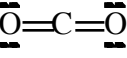


## Review of Molecular Shapes

Electron Areas	Bond Areas	Shape	Ball & Stick Model	Hybridization	Example
4	4	tetrahedral		 sp <sup>3</sup> hybridization	
4	3	trigonal pyramidal		 sp <sup>3</sup> hybridization	
4	2	angular		 sp <sup>3</sup> hybridization	
3	3	trigonal planar		 sp <sup>2</sup> hybridization	
3	2	angular		 sp <sup>2</sup> hybridization	
2	2	linear		 sp hybridization	

## Electronegativities

<b>F</b>	<b>O</b>	<b>Cl</b>	<b>N</b>	<b>Br</b>	<b>I</b>	<b>C</b>	<b>S</b>	<b>H</b>	<b>P</b>
4.0	3.4	3.2	3.0	2.7	2.6	2.6	2.2	2.2	2.2