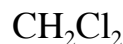
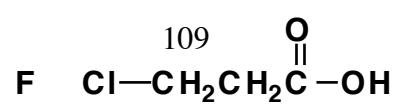
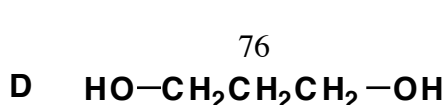
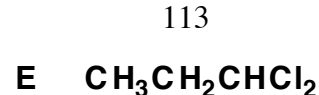
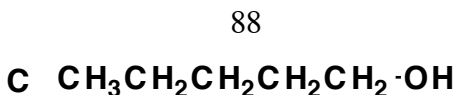
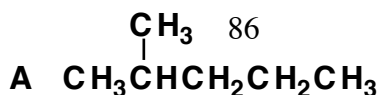
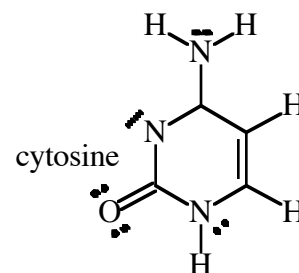
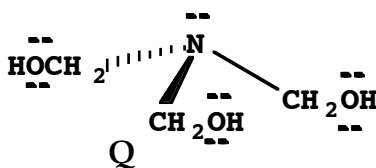


Which of these molecules are polar and which are nonpolar?



How many hydrogen bonding donor atoms and acceptor regions are in each of these molecules?



molecular weights above each structure

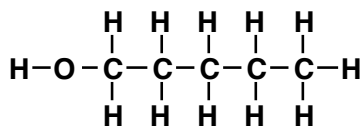
Which substance above, **E** or **F**, has the stronger intermolecular forces? Fill in table for comparison.

	London forces	Dipolar forces	H-bonding	Stronger overall?
E				
F				

Which substance above, **B** or **D**, will have the higher boiling point? Fill in table for comparison.

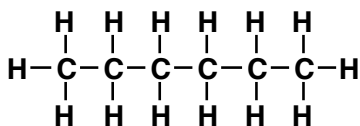
	London forces	Dipolar forces	H-bonding	Higher bp?
B				
D				

1-pentanol



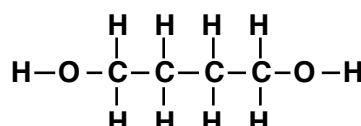
MW= 88 g/mole

n-hexane



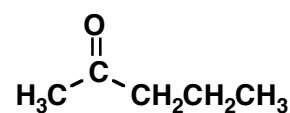
MW= 86 g/mole

1,4 butanediol



MW= 90 g/mole

2-pentanone



MW= 86 g/mole

Identify and compare the intermolecular attractions for each of the substances above.

Which of these substances has the highest boiling point?